Practice Worksheet: Names & Formulas of Acids

Review:

- When the anion does NOT contain Oxygen, & is a BINARY acid:
  Use the prefix *hydro* + **root of the anion's name** – *ic* + the word acid
  Examples: HCl = *hydro***chlor***ic* acid; HBr = *hydro***brom***ic* acid
- When the anion contains Oxygen, & is a TERNARY acid:
  The name depends on the name of the polyatomic anion. DON'T use the prefix '*hydro-*'! Examples: H_2SO_4 has the sulf**ate** anion, so the acid name will end in ***-ic:*** **Sulfuric acid.** H_2SO_3 has the sulf**ite** anion, so the name of the acid will end in ***-ous:*** **Sulfurous acid.**

  ATE → IC
  ITE → OUS

Write FORMULAS for the following:

Nitric acid	
Chloric acid	
Acetic acid	
Hydrobromic acid	
Sulfurous acid	
Chlorous acid	
Hydrochloric acid	
Phosphoric acid	
Nitrous acid	
Hydrofluoric acid	
Hypochlorous acid	
Hydroiodic acid	
Phosphorous acid	
Carbonic acid	
Perchloric acid	
Permanganic acid	
Sulfuric acid	
Hydrocyanic acid	

NAME the following :

$HClO$	
H_3PO_4	
HCl	
H_3BO_3	
H_2SO_4	
HNO_2	
HI	
$HC_2H_3O_2$	
HF	
H_3PO_3	
HCN	
$HClO_3$	
H_2CO_3	
H_2SO_3	
$HClO_2$	
HNO_3	
HBr	