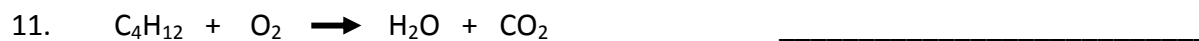
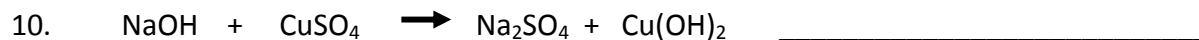
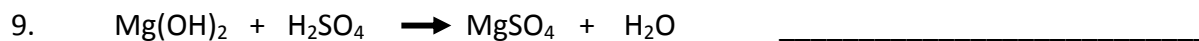
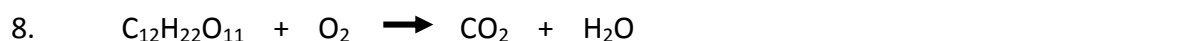


Classifying Chemical Reactions Worksheet

Name: _____ Period: _____

Classify each reaction as acid/base, redox, synthesis, decomposition, single replacement, double replacement or combustion. They may be more than one. The equations are not balanced.



- $\text{PbCl}_2 + \text{AgNO}_3 \rightarrow \text{Pb(NO}_3)_2 + \text{AgCl}$ double displacement
- $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$ synthesis / redox
- $\text{AlCl}_3 + \text{Na}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{NaCl}$ double displacement
- $\text{Zn} + \text{S} \rightarrow \text{ZnS}$ synthesis / redox
- $\text{Al}_2(\text{SO}_4)_3 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + \text{AlCl}_3$ double displacement
- $\text{Al}_2\text{S}_3 \rightarrow \text{Al} + \text{S}$ decomposition / redox
- $\text{H}_2\text{SO}_4 + \text{Fe} \rightarrow \text{H}_2 + \text{FeSO}_4$ single displacement / redox
- $\text{C}_{12}\text{H}_{22}\text{O}_{11} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ combustion / redox
- $\text{Mg(OH)}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + \text{H}_2\text{O}$ double displacement / acid base
- $\text{NaOH} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu(OH)}_2$ double displacement
- $\text{C}_4\text{H}_{12} + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{CO}_2$ combustion / redox
- $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ synthesis / redox
- $\text{Mg}_3(\text{PO}_4)_2 + \text{H}_2 \rightarrow \text{Mg} + \text{H}_3\text{PO}_4$ single displacement / redox
- $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$ decomposition / redox
- $\text{Cl}_2 + \text{KBr} \rightarrow \text{KCl} + \text{Br}_2$ single displacement / redox